Worksheet - Chemistry.

Name	Class ..X.....	Roll No.	Time 20 Min.	Max. Marks 10	Marks Obtained

TICK THE CORRECT OPTIONS:

1. **Which amongst the following represents a chemical change?** 1
   (*a*) Freezing of water
   (*b*) Liquefaction of air
   (*c*) Dissolution of sulphur in carbon disulphide
   (*d*) Cooking of food
2. **On heating magnesium ribbon in a bunsen flame, you will observe** 1
   (*a*) Magnesium ribbon burns with a yellow flame and changes into a white powder
   (*b*) Magnesium ribbon burns with a dazzling white flame and changes into a yellow powder
   (*c*) Magnesium ribbon burns with a dazzling white flame and changes into a white powder
   (*d*) Bluish flame is seen with the formation of white powder.
3. **The correct formula of Aluminium sulphate is** 1
   (*a*) Al_2SO_4 (*b*) $Al_3(SO_4)_2$
   (*c*) $AlSO_4$ (*d*) $Al_2(SO_4)_3$.
4. **The abbreviation used for a substance dissolved in water is** 1
   (*a*) au (*b*) as
   (*c*) aq (*d*) ao.
5. **On pouring dilute hydrochloric acid on zinc granules taken in a conical flask, which is the incorrect observation?** 1
   (*a*) Bubbles are seen rising from the surface of zinc granules
   (*b*) The evolved gas is hydrogen
   (*c*) The reaction mixture becomes hot
   (*d*) The gas evolved is chlorine
6. **A chemical equation is a symbolic representation of** 1
   (*a*) Chemical compound (*b*) Chemical reaction
   (*c*) Chemical element (*d*) None of these.
7. **What happens when dilute hydrochloric acid is added to iron filings?** 1
   (*a*) Hydrogen gas and iron chloride are produced
   (*b*) Chlorine gas and iron hydroxide are produced
   (*c*) No reaction takes place
   (*d*) Iron salt and water are produced
8. **The substances which undergo chemical change in a reaction are called** 1
   (*a*) Chemical compounds (*b*) Reactants
   (*c*) Products (*d*) Mixtures.
9. **Which amongst the following symbols is the correct representation for the formation of a precipitate?** 1
   (*a*) ↑ (*b*) ↓
   (*c*) → (*d*) ←
10. **Which of the following equation represents single displacement reaction?** 1
    (*a*) $NaCl\ (aq) + AgNO_3\ (aq) \longrightarrow NaNO_3\ (aq) + AgCl \downarrow$
    (*b*) $2Mg\ (s) + O_2\ (g) \longrightarrow 2MgO\ (s)$
    (*c*) $CaCO_3\ (s) \xrightarrow{\Delta} CaO\ (s) + CO_2\ (g)$
    (*d*) $Zn\ (s) + H_2SO_4\ (aq) \longrightarrow ZnSO_4\ (aq) + H_2\ (g)$

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TICK THE CORRECT OPTIONS:

1. **The method used to prevent rancidity of foods is** 1
   (*a*) Keep them in air tight containers (*b*) Add antioxidants to them
   (*c*) Flush them with nitrogen gas (*d*) All of these.

2. **Which amongst the following reactions represent a photochemical decomposition?** 1
   (*a*) $CaCO_3$ (s) $\longrightarrow$ CaO (s) + CO_2 (g) (*b*) 2AgBr (s) $\longrightarrow$ 2Ag (s) + Br_2 (g)
   (*c*) $2H_2O$ (l) $\longrightarrow$ $2H_2$ (g) + O_2 (g) (*d*) $2FeSO_4$ (s) $\longrightarrow$ Fe_2O_3 (s) + SO_2 (g) + SO_3 (g)

3. **Fatty foods becomes rancid due to the process of** 1
   (*a*) Oxidation (*b*) Reduction
   (*c*) Hydrogenation (*d*) Corrosion.

4. **In the reaction $2H_2S + SO_2 \longrightarrow 3S + 2H_2O$** 1
   (*a*) H_2S is getting oxidised (*b*) SO_2 is getting oxidised
   (*c*) H_2S is reduced (*d*) SO_2 is reducing agent.

5 **Balance the following equations:**

(*a*) $HNO_3 + Ca(OH)_2 \longrightarrow Ca(NO_3)_2 + H_2O$ 1

Ans. .....

(*b*) $NaOH + H_2SO_4 \longrightarrow Na_2SO_4 + H_2O$ 1

Ans. .....

(*c*) $NaCl + AgNO_3 \longrightarrow AgCl + NaNO_3$ 1

Ans. .....

6. **For carrying out decomposition reaction the energy required is in the form of** 1
   (*a*) Heat (*b*) Electricity
   (*c*) Light (*d*) All of these.

7 **On heating copper powder in a china dish, it gets coated with a black substance. This is due to** 1
   (*a*) Oxidation of copper (*b*) Reduction of copper
   (*c*) Corrosion of copper (*d*) None of these.

8 **Which of the following statements about the reaction below are incorrect?** 1

   2PbO + C (s) $\longrightarrow$ 2Pb (s) + CO_2 (g)

   (*i*) Lead is getting reduced (*ii*) Carbon dioxide is getting oxidised
   (*iii*) Carbon is getting oxidised (*iv*) Lead oxide is getting reduced
   (*a*) (*i*) and (*ii*) (*b*) (*i*) and (*iii*)
   (*c*) (*i*), (*ii*) and (*iii*) (*d*) (*iii*) and (*iv*)